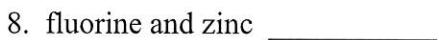
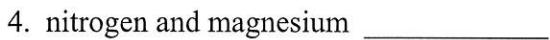
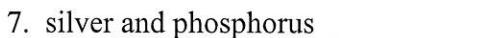
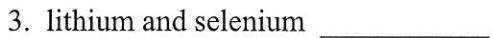
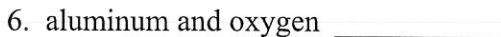
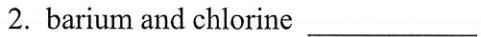
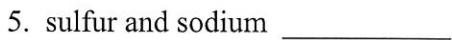
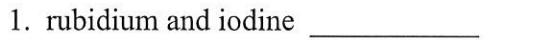
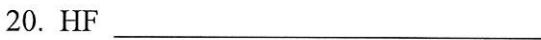
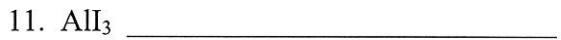


# Binary Ionic Compounds – Worksheet #1

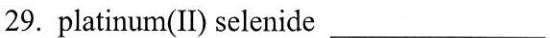
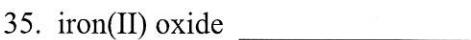
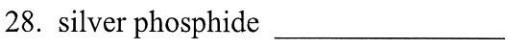
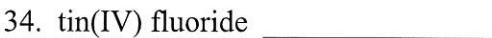
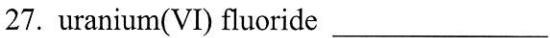
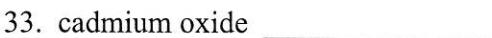
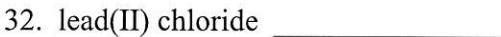
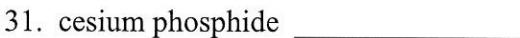
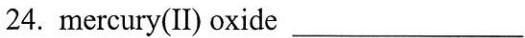
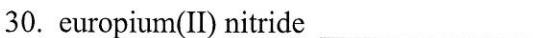
A. Write the formulas for the compounds formed from these elements. Remember, the cation is always written first.



B. Write the names for these binary ionic compounds.



C. Write the formulas for these binary ionic compounds.



## Binary Ionic Compounds – Worksheet #2

If the name of the compound is given, write the formula. If the formula of the compound is given, write the name.

1. KBr \_\_\_\_\_
2. V<sub>2</sub>O<sub>5</sub> \_\_\_\_\_
3. cobalt(III) oxide \_\_\_\_\_
4. barium phosphide \_\_\_\_\_
5. cadmium nitride \_\_\_\_\_
6. Cu<sub>3</sub>P \_\_\_\_\_
7. Ag<sub>2</sub>S \_\_\_\_\_
8. Sn<sub>3</sub>N<sub>4</sub> \_\_\_\_\_
9. radium iodide \_\_\_\_\_
10. beryllium selenide \_\_\_\_\_
11. Fe<sub>2</sub>S<sub>3</sub> \_\_\_\_\_
12. SrO \_\_\_\_\_
13. CrCl<sub>2</sub> \_\_\_\_\_
14. mercury(II) fluoride \_\_\_\_\_
15. lead(IV) bromide \_\_\_\_\_
16. CuSe \_\_\_\_\_
17. FeP \_\_\_\_\_
18. lithium oxide \_\_\_\_\_
19. cobalt(III) fluoride \_\_\_\_\_
20. CdI<sub>2</sub> \_\_\_\_\_