

Binary Ionic Compounds – Worksheet #1

A. Write the formulas for the compounds formed from these elements. Remember, the cation is always written first.

1. rubidium and iodine _____
2. barium and chlorine _____
3. lithium and selenium _____
4. nitrogen and magnesium _____
5. sulfur and sodium _____
6. aluminum and oxygen _____
7. silver and phosphorus _____
8. fluorine and zinc _____

B. Write the names for these binary ionic compounds.

9. Cs_2S _____
10. BaO _____
11. AlI_3 _____
12. MnO_2 _____
13. Tc_3P_4 _____
14. CdBr_2 _____
15. NaCl _____
16. FeF_3 _____
17. Mg_3N_2 _____
18. Ni_3P_2 _____
19. UO_2 _____
20. HF _____
21. CoN _____
22. K_2S _____

C. Write the formulas for these binary ionic compounds.

23. rubidium sulfide _____
24. mercury(II) oxide _____
25. calcium nitride _____
26. zinc bromide _____
27. uranium(VI) fluoride _____
28. silver phosphide _____
29. platinum(II) selenide _____
30. europium(II) nitride _____
31. cesium phosphide _____
32. lead(II) chloride _____
33. cadmium oxide _____
34. tin(IV) fluoride _____
35. iron(II) oxide _____
36. iron(III) oxide _____

Binary Ionic Compounds – Worksheet #2

If the name of the compound is given, write the formula. If the formula of the compound is given, write the name.

1. KBr _____
2. V_2O_5 _____
3. cobalt(III) oxide _____
4. barium phosphide _____
5. cadmium nitride _____
6. Cu_3P _____
7. Ag_2S _____
8. Sn_3N_4 _____
9. radium iodide _____
10. beryllium selenide _____
11. Fe_2S_3 _____
12. SrO _____
13. $CrCl_2$ _____
14. mercury(II) fluoride _____
15. lead(IV) bromide _____
16. CuSe _____
17. FeP _____
18. lithium oxide _____
19. cobalt(III) fluoride _____
20. CdI_2 _____