***Bonding!***

**Chemical Bond:**

The goal of all atoms is to be \_\_\_\_\_\_\_\_\_, \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, they will borrow, steal, or give to have 8.

**Covalent Bond:**

* Shared \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ are attracted to the nuclei of both atoms.
* \_\_\_\_\_\_\_\_\_\_\_\_\_\_ move back and forth between the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of each atom in the covalent bond.
* Each atom has a stable outer \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ some of the time.

**The neutral particle that is formed when atoms share electrons is called a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.**

Showing Covalent Bonds

1. To show a covalent bond, we begin with \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of the two elements.
2. Then combine the two dot notations.
3. This should show each element with a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
4. Then change the two dots to a \_\_\_\_\_\_\_\_\_\_\_\_ to represent the sharing of electrons.

**Covalent bonds and also known as \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.**

Label:

H2

*Practice: Show the bonding for two Chlorine, Cl, atoms:*

* + *Remember chlorine has 7 valence electrons.*

Types of Covalent Bonds:



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