

# ENDOTHERMIC VS, EXOTHERMIC REACTIONS

Research common reactions in nature or everyday use and label them as either Endothermic or Exothermic. Provide support for your decision (explain what energy is being used or produced). Each member of the group should provide at least one example not previously used. After research is completed, include all of the necessary information on the poster board.

Information to include:

- Common Source of Chemical Reaction (i.e. gas burning)
- Description of the area that the chemical reaction can be found (i.e. driving a car)
- Word Formula of Chemical Reaction (i.e. Octane and oxygen excited produces carbon dioxide and water)
- Equation Formula of Chemical Reaction ( $2\text{C}_8\text{H}_{18} + 25\text{O}_2 \rightarrow 16\text{CO}_2 + 18\text{H}_2\text{O}$ )
- Explanation of type of energy used or produced (Exothermic  $\rightarrow$  Heat and Light)
- Draw a picture (No example necessary)