

Ionic Compound Formula Writing Worksheet

Write chemical formulas for the compounds in each box. The names are found by finding the intersection between the cations and anions. Example: The first box is the intersection between the "zinc" cation and the "chloride" anion, so you should write "ZnCl₂", as shown.

	<i>zinc</i>	<i>iron (II)</i>	<i>iron (III)</i>	<i>gallium</i>	<i>silver</i>	<i>lead (IV)</i>
<i>chloride</i>	ZnCl ₂					
<i>acetate</i>						
<i>nitrate</i>						
<i>oxide</i>						
<i>nitride</i>						
<i>sulfate</i>						

Write the formulas for the following compounds:

- 1) copper (II) chloride _____
- 2) lithium acetate _____
- 3) vanadium (III) selenide _____
- 4) manganese (IV) nitride _____
- 5) beryllium oxide _____
- 6) sodium sulfate _____
- 7) aluminum arsenide _____
- 8) potassium permanganate _____
- 9) chromium (VI) cyanide _____
- 10) tin (II) sulfite _____
- 11) vanadium (V) fluoride _____
- 12) ammonium nitrate _____

Names & Formulas for Ionic Compounds

Give the name or formula of the following ionic compounds:

	Name		Formula
1)	Na_2CO_3 _____	21)	sodium phosphide _____
2)	NaOH _____	22)	magnesium nitrate _____
3)	MgBr_2 _____	23)	lead (II) sulfite _____
4)	KCl _____	24)	calcium phosphate _____
5)	FeCl_2 _____	25)	ammonium sulfate _____
6)	FeCl_3 _____	26)	silver cyanide _____
7)	Zn(OH)_2 _____	27)	aluminum sulfide _____
8)	Be_2SO_4 _____	28)	beryllium chloride _____
9)	CrF_2 _____	29)	copper (I) arsenide _____
10)	Al_2S_3 _____	30)	iron (III) oxide _____
11)	PbO _____	31)	gallium nitride _____
12)	Li_3PO_4 _____	32)	iron (II) bromide _____
13)	TiI_4 _____	33)	vanadium (V) phosphate _____
14)	Co_3N_2 _____	34)	calcium oxide _____
15)	Mg_3P_2 _____	35)	magnesium acetate _____
16)	$\text{Ga(NO}_2)_3$ _____	36)	aluminum sulfate _____
17)	Ag_2SO_3 _____	37)	copper (I) carbonate _____
18)	NH_4OH _____	38)	barium oxide _____
19)	Al(CN)_3 _____	39)	ammonium sulfite _____
20)	$\text{Be(CH}_3\text{COO)}_2$ _____	40)	silver bromide _____