**Molecular Nomenclature**

Compounds:

* Composed of more than one type of atom \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
* Separated \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ not \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
* No overall charge; they are electrically \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

Covalent Bonds

* Covalent bonds are formed between two \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
* Some non-metals bonds to an atom of the same element to form a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

*Mr.* B R I N Cl O F

Nomenclature of Molecules

* The naming of compounds is referred to as \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
* The nomenclature for molecular compounds is very easy if you know the meanings of the following Latin prefixes.












* The prefixes are used in the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of element’s name to match the number of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ per element.
* Then, add \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ to the second element.

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