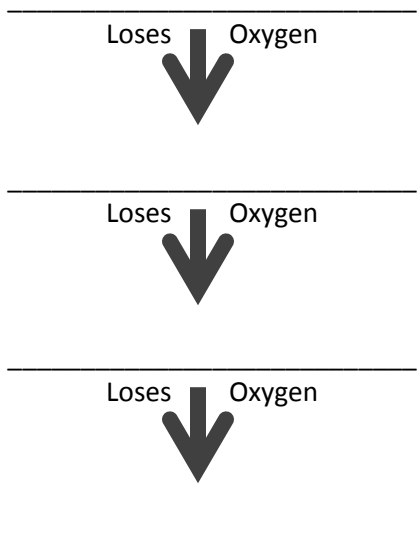


Naming Acids

Remember:

_____ is the only polyatomic cation that you need to know. Formula: _____

_____ is the only polyatomic anion with a 3- charge that you need to know. Formula: _____



Naming Acids:

_____	to →	_____
_____	to →	_____
_____	to →	_____
_____	to →	_____
_____	to →	_____

Examples:

_____ : Fluoride Ion	becomes →	_____ : HF
_____ : Nitrite Ion	becomes →	_____ : HNO ₂
_____ : Nitrate Ion	becomes →	_____ : _____
_____ : Phosphate Ion	becomes →	_____ : _____

Naming Acids

<i>Ion</i>	<i>Name</i>	<i>Acid Formula</i>	<i>Acid Name</i>
NH_4^+	Ammonium*		
NO_3^-	Nitrate	HNO_3	Nitric Acid
NO_2^-	Nitrite	HNO_2	Nitrous Acid
CN^-	Cyanide	HCN	Hydrocyanic Acid
SCN^-	Thiocyanate	HSCN	Thiocyanic Acid
ClO_4^-	Perchlorate*	HClO_4	Perchloric Acid
ClO_3^-	Chlorate*	HClO_3	Chloric Acid
ClO_2^-	Chlorite*	HClO_2	Chlorous Acid
ClO^-	Hypochlorite*	HClO	Hypochlorous Acid
$\text{C}_2\text{H}_3\text{O}_2^-$	Acetate	$\text{HC}_2\text{H}_3\text{O}_2$	Acetic Acid
MnO_4^-	Permanganate	HMnO_4	Permanganic Acid
SO_4^{2-}	Sulfate*	H_2SO_4	Sulfuric Acid
SO_3^{2-}	Sulfite*	H_2SO_3	Sulfurous Acid
HSO_4^-	Hydrogen sulfate or Bisulfate*	H_2SO_4	Sulfuric Acid
$\text{S}_2\text{O}_3^{2-}$	Thiosulfate	$\text{H}_2\text{S}_2\text{O}_3$	Thiosulfuric Acid
CO_3^{2-}	Carbonate	H_2CO_3	Carbonic Acid
HCO_3^-	Hydrogen carbonate or bicarbonate	H_2CO_3	Carbonic Acid
CrO_4^{2-}	Chromate	H_2CrO_4	Chromic Acid
$\text{Cr}_2\text{O}_7^{2-}$	Dichromate	$\text{H}_2\text{Cr}_2\text{O}_7$	
O_2^{2-}	Peroxide	H_2O_2	Hydrogen Peroxide (not really an acid)
$\text{C}_2\text{O}_4^{2-}$	Oxalate	$\text{H}_2\text{C}_2\text{O}_4$	Oxalic Acid
PO_4^{3-}	Phosphate*	H_3PO_4	Phosphoric Acid
HPO_4^{2-}	Hydrogen Phosphate*	H_3PO_4	Phosphoric Acid
H_2PO_4^-	Dihydrogen Phosphate*	H_3PO_4	Phosphoric Acid