**Naming Ionic Compounds**

**Name to Formula Steps**

1. Identify the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

 ex:

1. Switch the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ (drop the charge) to \_\_\_\_\_\_\_\_\_\_\_\_\_ of the other element.

 ex:

1. Simplify and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

 ex:

**Formula Rules**

* Simplify the number of elements:
	+ If the numbers are the same they \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
	+ If the numbers can be factored, \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

**Rules in Naming**

Cations: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Anions: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

* When combining the two the word “Ion” is dropped.

**Formula to Name**

1. Switch the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ to \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of the other element.

ex:

1. Label the first element as a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and the second as a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

ex:

1. Make sure the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of each is correct by looking at oxidation values.
2. Identify ions following rules for anions and cations.

ex:

**Identifying Names**

* If charges do not match the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of the element:
* Determine the needed value
* \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ both charges by that value
* If the cation is a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_with multiple possible \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, put the charge as

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ in parenthesis after the name.

**Naming Ionic Compounds**

**Name to Formula Steps**

1. Identify the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

ex:

1. Switch the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ (drop the charge) to \_\_\_\_\_\_\_\_\_\_\_\_\_ of the other element.

ex:

1. Simplify and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

ex:

**Formula Rules**

* Simplify the number of elements:
	+ If the numbers are the same they \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
	+ If the numbers can be factored, \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

**Rules in Naming**

Cations: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Anions: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

* When combining the two the word “Ion” is dropped.

**Formula to Name**

1. Switch the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ to \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of the other element.

ex:

1. Label the first element as a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and the second as a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

ex:

1. Make sure the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of each is correct by looking at oxidation values.
2. Identify ions following rules for anions and cations.

ex:

**Identifying Names**

* If charges do not match the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of the element:
* Determine the needed value
* \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ both charges by that value
* If the cation is a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_with multiple possible \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, put the charge as

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ in parenthesis after the name.