Name:	_ Date	_ Block		
Cups – Orbitals (subshells) Electrons – Beads White Beads – up arrows Blue Beads – down arrows	Orbital Notation Ch	emistry	Lab	
Put the correct number of electrons hold up to 2 electrons. Remember subshell individually before they wi	that the electrons wil	oshells ( I fill in a	cups) of the or n orbital which	bital. Each cup can has more than one
Important Information to remember				
EVERY subshell (line) can hold up	to electrons			
The s orbital can hold up to	electrons.			
The p orbital can hold up to	electrons.			
The d orbital can hold up to	electrons.			
The f orbital can hold up to	electrons.			
Fill in the chart below by writing in t	he orbital notation of	each el	ement. Use the	e cups and beads to
figure it out. Leave your beads in th	e cups until your tea	cher ha	s approved you	ur configuration. You
will do ONE configuration at a time.				ration before you move
on. (Electron configuration for neo	n looks like this: $1s^2$	$\frac{11 1}{2 s^2}$	1111 1111 11 11 11 11 11 11 11 11 11 11	Take special notice to

write the energy levels in large numbers, the orbital in lowercase letters, and the number of electrons

in superscript above the letter.)

		Teacher's
Element	Orbital Notation	initials
K		
F		
Sn		
Ra		
Ва		
Mn		
Fe		
Dy		
Ti		
As		
Kr		
Со		-
S		
Ni		

Si	
Eu	
Your	
choice p	
orbital	
Your	
choice d	
orbital	
Your	
choice f	
orbital	

Name:	.;	_ Date		Block		P	
Cups – Orbitals (subshell: Electrons – Beads White Beads – up : Blue Beads – down	arrows		otation Chen	nistry Lab			
Put the correct number of hold up to 2 electrons. Re subshell individually before	member	r that the e	into the subselectrons will f	hells (cups) ill in an orbit	of the orbital	. Each cup of more than c	can
Important Information to re	emembe	er:					
EVERY subshell (line) car	n hold up	o to	_ electrons.				
The s orbital can hold up	to	electr	rons.				
The p orbital can hold up	to	electr	rons.				
The d orbital can hold up	to	electr	rons.				
The f orbital can hold up to	0	elec	ctrons.				
Fill in the chart below by v	vriting in	the orbital	l notation of e	ach elemen	t. Use the cu <sub>l</sub>	ps and bead	ls to
figure it out. Leave your be	eads in t	he cups u	ntil vour teach	er has appr	oved vour co	nfiguration	You

Fill in the chart below by writing in the orbital notation of each element. Use the cups and beads to figure it out. Leave your beads in the cups until your teacher has approved your configuration. You will do ONE configuration at a time. Be sure to get approval after each configuration before you move on. (Electron configuration for neon looks like this:  $\frac{1}{1}$   $\frac{1}{2}$   $\frac{1}{2$ 

C:		
Si		
F		
Eu		
Your		
choice p	A 9	
orbital		
Your		
		[Sep
choice d		
orbital		
Your		
choice f		
1.76		
orbital		

* * * * * * * * * * * * * * * * * * * *				
Name:	_ Date	Block		
Cups – Orbitals (subshells) Electrons – Beads White Beads – up arrows Blue Beads – down arrows	Orbital Notation Cl	nemistry Lab		
Put the correct number of electron hold up to 2 electrons. Remember subshell individually before they w	that the electrons w	ubshells (cups) of the orb rill fill in an orbital which	oital. Each o has more th	up can an one
Important Information to remembe	r:			
EVERY subshell (line) can hold up	to electrons	S.	*	
The s orbital can hold up to	electrons.			
The p orbital can hold up to	electrons.	:		
The d orbital can hold up to	electrons.			
The f orbital can hold up to	electrons.			
Fill in the chart below by writing in	the orbital notation o	of each element. Use the	cups and b	eads to

Fill in the chart below by writing in the orbital notation of each element. Use the cups and beads to figure it out. Leave your beads in the cups until your teacher has approved your configuration. You will do ONE configuration at a time. Be sure to get approval after each configuration before you move on. (Electron configuration for neon looks like this:  $\frac{1}{1} \cdot \frac{1}{2} \cdot \frac{1}{2$ 

f O:		γ
Si		
		A
Eu		
Eu		
2		
Your		
choice p		
choice b		
was many sal		
orbital		
Your		
Toul		
choice d	a a contract of the contract o	
orbital		
Orbital		
Your		
oboice f		
choice f		
		la la
orbital		
)	*	