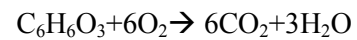


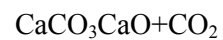
Name:

## Percent Yield Practice

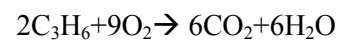
1. For the balanced equation shown below, if the reaction of 40.8 grams of  $C_6H_6O_3$  produces a 39.0% yield, how many grams of  $H_2O$  would be produced?



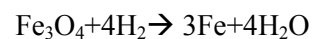
2. For the balanced equation shown below, if the reaction of 20.7 grams of  $CaCO_3$  produces 6.81 grams of  $CaO$ , what is the percent yield?



3. For the balanced equation shown below, if the reaction of 91.3 grams of  $C_3H_6$  produces a 81.3% yield, how many grams of  $CO_2$  would be produced?



4. For the balanced equation shown below, if the reaction of 0.112 grams of  $H_2$  produces 0.745 grams of  $H_2O$ , what is the percent yield?



5. For the balanced equation shown below, if the reaction of 77.0 grams of  $\text{CaCN}_2$  produces 27.1 grams of  $\text{NH}_3$ , what is the percent yield?

