

Name:

### Unit 4 Study Guide Part 2

1. In reactions atoms are not lost but \_\_\_\_\_.
2. \_\_\_\_\_ yields \_\_\_\_\_ in a chemical reaction.
3. Subscripts provide a \_\_\_\_\_ between the number of \_\_\_\_\_ for each element involved.
4. Energy is stored in \_\_\_\_\_ in a chemical reaction.
5. A mole is the \_\_\_\_\_ of a substance.
6. \_\_\_\_\_ is defined as the percentage by mass of each element in a compound
7. Label the following letters with the word that they represent. Explain why it is important to use these within a chemical reaction.
  - a. (l)
  - b. (s)
  - c. (g)
  - d. (aq)
8. What is a Chemical Reaction:
9. What are three indications of a chemical reaction?
10. Explain how the conservation of matter relates to chemical reactions.
11. Explain (not list) the five different chemical equations thus far.

12. Compare and contrast word formulas and equation formulas.

13. Describe the steps to balance an equation, use a minimum of four steps.

14. Compare and contrast reactions being described as either endothermic or exothermic.

15. Compare and contrast formula mass to molar mass.